

8.15 STANDARDNÍ ELEKTRODOVÉ POTENCIÁLY E° PŘI TEPLOTĚ 25 °C
A KOEFICIENTY JEJICH TEPLOTNÍ ZÁVISLOSTI dE°/dT

Použité zkratky:

Ac = acetát

Phen = fenantrolin

Cd(Hg) = amalgám kadmia

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|---|----------------|-------------------------------------|
| $\text{Ag}^+ + e^- = \text{Ag}$ | 0,799 6 | -1,000 |
| $\text{AgAc} + e^- = \text{Ag} + \text{Ac}^-$ | 0,643 | — |
| $\text{AgBr} + e^- = \text{Ag} + \text{Br}^-$ | 0,071 33 | — |
| $\text{Ag}_2\text{C}_2\text{O}_4 + 2e^- = 2\text{Ag} + \text{C}_2\text{O}_4^{2-}$ | 0,464 7 | — |
| $\text{AgCl} + e^- = \text{Ag} + \text{Cl}^-$ | 0,222 2 | -0,658 |
| $\text{AgI} + e^- = \text{Ag} + \text{I}^-$ | -0,152 24 | — |
| $2\text{AgO} + \text{H}_2 + 2e^- = \text{Ag}_2\text{O} + 2\text{OH}^-$ | 0,607 | -1,117 |
| $\text{Ag}_2\text{O} + \text{H}_2\text{O} + 2e^- = 2\text{Ag} + 2\text{OH}^-$ | 0,342 | -1,337 |
| $\text{Ag}_2\text{SO}_4 + 2e^- = 2\text{Ag} + \text{SO}_4^{2-}$ | 0,654 | — |
| $\text{Ag}_2\text{S} + 2e^- = 2\text{Ag} + \text{S}^{2-}$ | -0,691 | — |
| $\text{Al}^{3+} + 3e^- = \text{Al}$ | -1,662 | 0,504 |
| $\text{H}_2\text{AlO}_3^- + \text{H}_2\text{O} + 3e^- = \text{Al} + 4\text{OH}^-$ | -2,33 | -0,93 |
| $\text{As} + 3\text{H}^+ + 3e^- = \text{AsH}_3$ | -0,608 | -0,05 |
| $\text{HAsO}_2 + 3\text{H}^+ + 3e^- = \text{As} + 2\text{H}_2\text{O}$ | 0,248 | -0,510 |
| $\text{Au}^+ + e^- = \text{Au}$ | 1,692 | — |
| $\text{Au}^{3+} + 2e^- = \text{Au}^+$ | 1,401 | — |
| $\text{Au}^{3+} + 3e^- = \text{Au}$ | 1,498 | — |
| $\text{Au(OH)}_3 + 3\text{H}^+ + 3e^- = \text{Au} + 3\text{H}_2\text{O}$ | 1,45 | -0,206 |
| $\text{AuCl}_4^- + 3e^- = \text{Au} + 4\text{Cl}^-$ | 1,002 | — |
| $\text{H}_2\text{BO}_3^- + 5\text{H}_2\text{O} + 8e^- = \text{BH}_4^- + 8\text{OH}^-$ | -1,24 | — |
| $\text{H}_2\text{BO}_3^- + \text{H}_2\text{O} + 3e^- = \text{B} + 4\text{OH}^-$ | -1,79 | -1,147 |
| $\text{H}_3\text{BO}_3 + 3\text{H}^+ + 3e^- = \text{B} + 3\text{H}_2\text{O}$ | -0,869 8 | -0,481 |
| $\text{Ba}^{2+} + 2e^- = \text{Ba}$ | -2,912 | -0,395 |
| $\text{Ba}^{2+} + 2e^- = \text{Ba(Hg)}$ | -1,570 | — |
| $\text{Ba(OH)}_2 + 2e^- = \text{Ba} + 2\text{OH}^-$ | -2,99 | -0,93 |
| $\text{Be}^{2+} + 2e^- = \text{Be}$ | -1,847 | 0,565 |
| $\text{Bi}_2\text{O}_3 + 3\text{H}_2\text{O} + 6e^- = 2\text{Bi} + 6\text{OH}^-$ | -0,46 | -1,214 |
| $\text{BiO}^+ + 2\text{H}^+ + 3e^- = \text{Bi} + \text{H}_2\text{O}$ | 0,320 | — |
| $\text{Br}_2(\text{aq}) + 2e^- = 2\text{Br}^-$ | 1,087 3 | -0,478 |
| $\text{Br}_2(\text{l}) + 2e^- = 2\text{Br}^-$ | 1,066 | -0,629 |
| $\text{BrO}_3^- + 3\text{H}_2\text{O} + 6e^- = \text{Br}^- + 6\text{OH}^-$ | 0,61 | -1,287 |
| $\text{BrO}_3^- + 6\text{H}^+ + 6e^- = \text{Br}^- + 3\text{H}_2\text{O}$ | 1,423 | — |

pokračování 1

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|--|----------------|-------------------------------------|
| $\text{C} + 4\text{H}^+ + 4e^- = \text{CH}_4$ | 0,131 6 | -0,209 |
| $\text{CO} + 2\text{H}^+ + 2e^- = \text{C} + \text{H}_2\text{O}$ | 0,517 8 | -1,310 |
| $\text{CO}_2 + 4\text{H}^+ + 4e^- = \text{C} + 2\text{H}_2\text{O}$ | 0,207 3 | -0,853 |
| $\text{CO}_3^{2-} + 3\text{H}_2\text{O} + 4e^- = \text{C} + 6\text{OH}^-$ | -0,766 7 | -1,232 |
| $\text{Ca}^{2+} + 2e^- = \text{Ca}$ | -2,868 | -0,175 |
| $\text{Ca(OH)}_2 + 2e^- = \text{Ca} + 2\text{OH}^-$ | -3,02 | -0,965 |
| $\text{Cd}^{2+} + 2e^- = \text{Cd}$ | -0,403 0 | -0,093 |
| $\text{Cd}^{2+} + 2e^- = \text{Cd(Hg)}$ | -0,352 1 | — |
| $\text{Cd(OH)}_2 + 2e^- = \text{Cd(Hg)} + 2\text{OH}^-$ | -0,809 | -1,014 |
| $\text{Ce}^{3+} + 3e^- = \text{Ce}$ | -2,483 | 0,101 |
| $\text{Ce}^{3+} + 3e^- = \text{Ce(Hg)}$ | -1,437 3 | — |
| $\text{Ce}^{4+} + e^- = \text{Ce}^{3+}$ | 1,61 | — |
| $\text{Cl}_2 + 2e^- = 2\text{Cl}^-$ | 1,358 3 | -1,260 |
| $\text{HClO} + \text{H}^+ + e^- = \frac{1}{2}\text{Cl}^- + \text{H}_2\text{O}$ | 1,611 | -0,14 |
| $\text{ClO}^- + \text{H}_2\text{O} + 2e^- = \text{Cl}^- + 2\text{OH}^-$ | 0,89 | -1,079 |
| $\text{HClO}_2 + 2\text{H}^+ + 2e^- = \text{HClO} + \text{H}_2\text{O}$ | 1,645 | -0,55 |
| $\text{ClO}_2^- + 2\text{H}_2\text{O} + 4e^- = \text{Cl}^- + 2\text{OH}^-$ | 0,76 | — |
| $\text{ClO}_3^- + 6\text{H}^+ + 6e^- = \text{Cl}^- + 3\text{H}_2\text{O}$ | 1,451 | — |
| $\text{ClO}_3^- + 3\text{H}_2\text{O} + 6e^- = \text{Cl}^- + 6\text{OH}^-$ | 0,620 | — |
| $\text{ClO}_4^- + 8\text{H}^+ + 8e^- = \text{Cl}^- + 4\text{H}_2\text{O}$ | 1,389 | — |
| $\text{Co}^{2+} + 2e^- = \text{Co}$ | -0,280 | 0,06 |
| $[\text{Co(NH}_3)_6]^{3+} + e^- = [\text{Co(NH}_3)_6]^{2+}$ | 0,108 | — |
| $\text{Co(OH)}_2 + 2e^- = \text{Co} + 2\text{OH}^-$ | -0,730 | -1,064 |
| $\text{Co(OH)}_3 + e^- = \text{Co(OH)}_2 + \text{OH}^-$ | 0,170 | -0,80 |
| $\text{Cr}^{2+} + 2e^- = \text{Cr}$ | -0,913 | — |
| $\text{Cr}^{3+} + e^- = \text{Cr}^{2+}$ | -0,407 | — |
| $\text{Cr}^{3+} + 3e^- = \text{Cr}$ | -0,744 | 0,468 |
| $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6e^- = 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ | 1,33 | -1,263 |
| $\text{Cr(OH)}_3 + 3e^- = \text{Cr} + 3\text{OH}^-$ | -1,48 | — |
| $\text{Cs}^+ + e^- = \text{Cs}$ | -2,92 | -1,197 |
| $\text{Cu}^+ + e^- = \text{Cu}$ | 0,521 | -0,058 |
| $\text{Cu}^{2+} + e^- = \text{Cu}^+$ | 0,153 | — |
| $\text{Cu}^{2+} + 2e^- = \text{Cu}$ | 0,431 9 | -0,058 |
| $\text{Cu}^{2+} + 2e^- = \text{Cu(Hg)}$ | 0,345 | — |
| $\text{Cu}_2\text{O} + \text{H}_2\text{O} + 2e^- = 2\text{Cu} + 2\text{OH}^-$ | -0,360 | -1,326 |
| $\text{Cu(OH)}_2 + 2e^- = \text{Cu} + 2\text{OH}^-$ | -0,080 | -0,725 |
| $\text{D}^+ + e^- = \frac{1}{2}\text{D}_2$ | -0,003 4 | — |
| $2\text{D}^+ + 2e^- = \text{D}_2$ | -0,044 | — |
| $\text{Eu}^{2+} + 2e^- = \text{Eu}$ | -3,395 | — |
| $\text{Eu}^{3+} + 3e^- = \text{Eu}$ | -2,407 | — |

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|--|----------------|-------------------------------------|
| $F_2 + 2H^+ + 2e^- = 2HF$ | 3,06 | -0,60 |
| $F_2 + 2e^- = 2F^-$ | 2,866 | -1,830 |
| $F_2O + 2H^+ + 4e^- = H_2O + 2F^-$ | 2,153 | -1,184 |
| $Fe^{2+} + 2e^- = Fe$ | -0,441 | 0,052 |
| $Fe^{3+} + 3e^- = Fe$ | -0,037 | — |
| $Fe^{3+} + e^- = Fe^{2+}$ | 0,771 | 1,188 |
| $[Fe(CN)_6]^{3+} + e^- = [Fe(CN)_6]^{4+}$ | 0,358 | — |
| $FeO_4^{2-} + 8H^+ + 3e^- = Fe^{3+} + 4H_2O$ | 2,2 | -0,85 |
| $Fe(OH)_3 + e^- = Fe(OH)_2 + OH^-$ | -0,56 | -0,96 |
| $Fe(OH)_2 + 2e^- = Fe + 2OH^-$ | -0,877 | -1,06 |
| $[Fe(Phen)_3]^{3+} + e^- = [Fe(Phen)_3]^{2+}$ | 1,147 | — |
| [ferricinium] ⁺ + e ⁻ = ferrocen | 0,400 | — |
| $Ga^{3+} + 3e^- = Ga$ | -0,529 | 0,67 |
| $H_2GaO_3 + H_2O + 3e^- = Ga + 4OH^-$ | -1,219 | — |
| $Ge^{2+} + 2e^- = Ge$ | 0,24 | — |
| $Ge^{4+} + 4e^- = Ge$ | 0,124 | — |
| $GeO_2 + 2H^+ + 2e^- = GeO + H_2O$ | -0,118 | — |
| $GeO_2 + 4H^+ + 4e^- = Ge + 2H_2O$ | -0,15 | -0,335 |
| $2H^+ + 2e^- = H_2$ | 0,000 0 | 0,000 |
| $H_2 + 2e^- = 2H^-$ | -2,23 | -1,57 |
| $2H_2O + 2e^- = H_2 + 2OH^-$ | -0,827 7 | -0,834 2 |
| $H_2O_2 + 2H^+ + 2e^- = 2H_2O$ | 1,776 | — |
| $Hg^{2+} + 2e^- = Hg$ | 0,788 | — |
| $2Hg^{2+} + 2e^- = Hg_2^{2+}$ | 0,920 | — |
| $Hg_2^{2+} + 2e^- = 2Hg$ | 0,797 3 | — |
| $Hg_2(Ac)_2 + 2e^- = 2Hg + 2Ac^-$ | 0,511 63 | — |
| $Hg_2Br_2 + 2e^- = 2Hg + 2Br^-$ | 0,139 2 | — |
| $Hg_2Cl_2 + 2e^- = 2Hg + 2Cl^-$ | 0,267 6 | -0,317 |
| $Hg_2I_2 + 2e^- = 2Hg + 2I^-$ | -0,040 5 | — |
| $Hg_2O + H_2O + 2e^- = 2Hg + 2OH^-$ | 0,123 | — |
| $HgO + H_2O + 2e^- = Hg + 2OH^-$ | 0,097 7 | -1,120 |
| $Hg_2SO_4 + 2e^- = 2Hg + SO_4^{2-}$ | 0,612 5 | — |
| $I_2 + 2e^- = 2I^-$ | 0,535 5 | -0,148 |
| $I_3^- + 2e^- = 3I^-$ | 0,536 | — |
| $2IO_3^- + 12H^+ + 10e^- = I_2 + 6H_2O$ | 1,195 | -0,364 |
| $IO^- + H_2O + 2e^- = I^- + 2OH^-$ | 0,485 | — |
| $2HIO + 2H^+ + 2e^- = I_2 + 2H_2O$ | 1,439 | 0,42 |
| $H_3IO_6 + H^+ + 2e^- = IO_3^- + 3H_2O$ | 1,601 | — |
| $In^+ + e^- = In$ | -0,14 | — |
| $In^{3+} + 3e^- = In$ | -0,338 2 | 0,40 |

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|---|----------------|-------------------------------------|
| $Ir^{3+} + 3e^- = Ir$ | 1,156 | — |
| $Ir_2O_3 + 3H_2O + 6e^- = 2Ir + 6OH^-$ | 0,098 | — |
| $K^+ + e^- = K$ | -2,931 | — |
| $La^{3+} + 3e^- = La$ | -2,522 | — |
| $La(OH)_3 + 3e^- = La + 3OH^-$ | -2,90 | -0,95 |
| $Li^+ + e^- = Li$ | -3,040 1 | -0,534 |
| $Mg^+ + e^- = Mg$ | -2,70 | — |
| $Mg^{2+} + 2e^- = Mg$ | -2,372 | 0,103 |
| $Mg(OH)_2 + 2e^- = Mg + 2OH^-$ | -2,690 | -0,945 |
| $Mn^{2+} + 2e^- = Mn$ | -1,185 | -0,08 |
| $Mn^{3+} + 3e^- = Mn^{2+}$ | 1,541 5 | — |
| $MnO_2 + 4H^+ + 2e^- = Mn^{2+} + 2H_2O$ | 1,224 | -0,661 |
| $MnO_4^- + 4H^+ + 3e^- = MnO_2 + 2H_2O$ | 1,679 | -0,666 |
| $MnO_4^- + 8H^+ + 5e^- = Mn^{2+} + 4H_2O$ | 1,507 | -0,666 |
| $Mn(OH)_2 + 2e^- = Mn + 2OH^-$ | -1,56 | -1,079 |
| $Mn(OH)_3 + e^- = Mn(OH)_2 + OH^-$ | 0,15 | -0,903 |
| $Mo^{3+} + 3e^- = Mo$ | -0,200 | — |
| $N_2 + 2H_2O + 6H^+ + 6e^- = 2NH_4OH$ | 0,092 | — |
| $N_2 + 8H^+ + 6e^- = 2NH_4^+$ | 0,274 6 | -0,618 |
| $N_2O + 2H^+ + 2e^- = N_2 + H_2O$ | 1,766 | — |
| $N_2O_4 + 2e^- = 2NO_2^-$ | 0,867 | — |
| $N_2O_4 + 2H^+ + 2e^- = 2HNO_2$ | 1,065 | — |
| $N_2O_4 + H^+ + 4e^- = 2NO + 2H_2$ | 1,035 | — |
| $2NO + 2H^+ + 2e^- = N_2O + H_2O$ | 1,591 | — |
| $HNO_2 + H^+ + e^- = NO + H_2O$ | 0,983 | — |
| $NO_3^- + 3H^+ + 2e^- = HNO_2 + H_2O$ | 0,934 | — |
| $NO_3^- + 4H^+ + 3e^- = NO + 2H_2O$ | 0,957 | — |
| $2NO_3^- + 4H^+ + 2e^- = N_2O_4 + 2H_2O$ | 0,803 | — |
| $NO_3^- + H_2O + 2e^- = NO_2^- + 2OH^-$ | 0,01 | — |
| $Na^+ + e^- = Na$ | -2,714 | -0,772 |
| $Nb^{3+} + 3e^- = Nb$ | -1,099 | — |
| $Nb_2O_5 + 10H^+ + 10e^- = 2Nb + 5H_2O$ | -0,644 | -0,39 |
| $Nd^{3+} + 3e^- = Nd$ | -2,431 | — |
| $Ni^{2+} + 2e^- = Ni$ | -0,257 | 0,06 |
| $Ni(OH)_2 + 2e^- = Ni + 2OH^-$ | -0,72 | -1,04 |
| $NiO_2 + 2H_2O + 2e^- = Ni(OH)_2 + 2OH^-$ | 0,490 | — |
| $O_2 + 2H^+ + 2e^- = H_2O_2$ | 0,695 | -1,033 |
| $O_2 + 4H^+ + 4e^- = 2H_2O$ | 1,229 | -0,846 |
| $O_2 + 2H_2O + 2e^- = H_2O_2 + 2OH^-$ | -0,146 | — |
| $O_2 + 2H_2O + 4e^- = 4OH^-$ | 0,401 | -1,680 |

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|--|----------------|-------------------------------------|
| $O_3 + 2H^+ + 2e^- = O_2 + H_2O$ | 2,076 | — |
| $O_3 + H_2O + 2e^- = O_2 + 2OH^-$ | 1,24 | — |
| $P + 3H^+ + 3e^- = PH_3$ | -0,063 | -0,104 |
| $H_2PO_2^- + e^- = P + 2OH^-$ | -1,82 | — |
| $H_3PO_2 + H^+ + 3e^- = P + 2H_2O$ | -0,508 | — |
| $H_3PO_3 + 2H^+ + 2e^- = H_3PO_2 + H_2O$ | -0,499 | — |
| $H_3PO_3 + 3H^+ + 3e^- = P + 3H_2O$ | -0,454 | -0,36 |
| $H_3PO_4 + 2H^+ + 2e^- = H_3PO_3 + H_2O$ | -0,276 | -0,36 |
| $PO_4^{3-} + 2H_2O + 2e^- = HPO_3^{2-} + 3OH^-$ | -1,05 | — |
| $Pb^{2+} + 2e^- = Pb$ | -0,126 2 | -0,451 |
| $Pb^{2+} + 2e^- = Pb(Hg)$ | -0,120 5 | — |
| $PbBr_2 + 2e^- = Pb + 2Br^-$ | -0,284 | — |
| $PbCl_2 + 2e^- = Pb + 2Cl^-$ | -0,267 5 | — |
| $PbO + H_2O + 2e^- = Pb + 2OH^-$ | -0,580 | — |
| $PbO_2 + 4H^+ + 2e^- = Pb^{2+} + 2H_2O$ | 1,455 | -0,238 |
| $PbO_2 + H_2O + 2e^- = PbO + 2OH^-$ | 0,247 | -1,194 |
| $PbO_2 + SO_4^{2-} + 4H^+ + 2e^- = PbSO_4 + 2H_2O$ | 1,691 3 | 0,326 |
| $PbSO_4 + 2e^- = Pb + SO_4^{2-}$ | -0,358 8 | — |
| $Pd^{2+} + 2e^- = Pd$ | 0,987 | — |
| $[PdCl_4]^{2-} + 2e^- = Pd + 4Cl^-$ | 0,591 | — |
| $[PdCl_6]^{2-} + 2e^- = [PdCl_4]^{2-} + 2Cl^-$ | 1,288 | — |
| $Pt^{2+} + 2e^- = Pt$ | 1,118 | — |
| $[PtCl_4]^{2-} + 2e^- = Pt + 4Cl^-$ | 0,755 | — |
| $[PtCl_6]^{2-} + 2e^- = [PtCl_4]^{2-} + 2Cl^-$ | 0,68 | — |
| $Pt(OH)_2 + 2e^- = Pt + 2OH^-$ | 0,15 | -1,144 |
| $Pt(OH)_2 + 2H^+ + 2e^- = Pt + 2H_2O$ | 0,98 | -0,310 |
| $Ra^{2+} + 2e^- = Ra$ | -2,916 | -0,59 |
| $Rb^+ + e^- = Rb$ | -2,925 | -1,245 |
| $Re^{3+} + 3e^- = Re$ | 0,300 | — |
| $ReO_4^- + 4H^+ + 3e^- = ReO_2 + 2H_2O$ | 0,510 | — |
| $ReO_2 + 4H^+ + 4e^- = Re + 2H_2O$ | 0,251 3 | — |
| $ReO_4^- + 8H^+ + 7e^- = Re + 4H_2O$ | 0,362 | -0,51 |
| $Rh^+ + e^- = Rh$ | 0,600 | — |
| $Rh^{3+} + 3e^- = Rh$ | 0,788 | — |
| $[RhCl_6]^{3-} + 3e^- = Rh + 6Cl^-$ | 0,431 | — |
| $Ru^{2+} + 2e^- = Ru$ | 0,455 | — |
| $Ru^{3+} + e^- = Ru^{2+}$ | 0,248 7 | — |
| $RuO_2 + 4H^+ + 2e^- = Ru^{2+} + 2H_2O$ | 1,120 | — |
| $S + 2e^- = S^{2-}$ | -0,476 3 | — |
| $S + 2H^+ + 2e^- = H_2S(aq)$ | 0,142 | 0,209 |

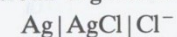
| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|---|----------------|-------------------------------------|
| $S + H_2O + 2e^- = HS^- + OH^-$ | -0,478 | — |
| $S_2O_6^{2-} + 4H^+ + 2e^- = 2H_2SO_3$ | 0,564 | — |
| $S_2O_8^{2-} + 2e^- = 2SO_4^{2-}$ | 2,010 | -1,26 |
| $H_2SO_3 + 4H^+ + 4e^- = 2S + 3H_2O$ | 0,449 | -0,66 |
| $SO_4^{2-} + 4H^+ + 2e^- = H_2SO_3 + H_2O$ | 0,172 | — |
| $SO_4^{2-} + H_2O + 2e^- = SO_3^{2-} + 2OH^-$ | -0,93 | — |
| $Sb + 3H^+ + 3e^- = SbH_3$ | -0,510 | -0,06 |
| $Sb_2O_3 + 6H^+ + 6e^- = 2Sb + 3H_2O$ | 0,152 | -0,375 |
| $Se^{3+} + 3e^- = Se$ | -2,077 | 0,25 |
| $Se + 2e^- = Se^{2-}$ | -0,924 | -0,89 |
| $Se + 2H^+ + 2e^- = H_2Se(aq)$ | -0,399 | -0,028 |
| $H_2SeO_3 + 4H^+ + 4e^- = Se + 3H_2O$ | 0,740 | -0,250 |
| $SeO_3^{2-} + 3H_2O + 4e^- = Se + 6OH^-$ | -0,366 | -1,318 |
| $SeO_4^{2-} + H_2O + 2e^- = SeO_3^{2-} + 2OH^-$ | 0,05 | -1,187 |
| $SiF_6^{2-} + 4e^- = Si + 6F^-$ | -1,24 | — |
| $SiO_3^{2-} + 3H_2O + 4e^- = Si + 6OH^-$ | -1,697 | — |
| $SiO_2 + 4H^+ + 4e^- = Si + 2H_2O$ | -0,857 | -0,374 |
| $Si + 4H^+ + 4e^- = SiH_4$ | 0,102 | -0,197 |
| $Sn^{2+} + 2e^- = Sn$ | -0,136 | -0,282 |
| $Sn^{4+} + 2e^- = Sn^{2+}$ | 0,151 | — |
| $HSnO_2^- + H_2O + 2e^- = Sn + 2OH^-$ | -0,909 | — |
| $Sr^{2+} + 2e^- = Sr$ | -2,888 | -0,191 |
| $Sr^{2+} + 2e^- = Sr(Hg)$ | -1,793 | — |
| $Sr(OH)_2 + 2e^- = Sr + 2OH^-$ | -2,88 | -0,96 |
| $Ta_2O_5 + 10H^+ + 10e^- = 5Ta + 5H_2O$ | -0,750 | -0,377 |
| $Tc^{2+} + 2e^- = Tc$ | 0,400 | — |
| $TcO_4^- + 4H^+ + 3e^- = TcO_2 + 2H_2O$ | 0,782 | — |
| $Te + 2e^- = Te^{2-}$ | -1,143 | — |
| $Te + 2H^+ + 2e^- = H_2Te$ | -0,718 | 0,280 |
| $TeO_2 + 4H^+ + 4e^- = Te + 2H_2O$ | 0,593 | -0,370 |
| $H_6TeO_6 + 2H^+ + 2e^- = TeO_2 + 4H_2O$ | 1,02 | 0,13 |
| $Th^{4+} + 4e^- = Th$ | -1,899 | — |
| $ThO_2 + 4H^+ + 4e^- = Th + 2H_2O$ | -1,789 | — |
| $Ti^{2+} + 2e^- = Ti$ | -1,628 | — |
| $Ti^{3+} + e^- = Ti^{2+}$ | -0,368 | — |
| $Tl^+ + e^- = Tl$ | -0,336 6 | -1,327 |
| $Tl^+ + e^- = Tl(Hg)$ | -0,333 8 | — |
| $Tl^{3+} + 2e^- = Tl^+$ | 1,252 | 0,89 |
| $TlCl + e^- = Tl + Cl^-$ | -0,556 8 | — |
| $TlOH + e^- = Tl + OH^-$ | -0,34 | — |

| Elektrodová reakce | E° V | dE°/dT mV K ⁻¹ |
|--|----------------|-------------------------------------|
| $U^{3+} + 3e^- = U$ | -1,798 | — |
| $U^{4+} + e^- = U^{3+}$ | -0,607 | — |
| $UO_2 + 2H_2O + 4e^- = U + 4OH^-$ | -2,39 | -1,220 |
| $UO_4^{2-} + 4H_2O + 2e^- = U(OH)_4 + 4OH^-$ | -1,618 | — |
| $UO_2^{2+} + 4H^+ + 2e^- = U^{4+} + 2H_2O$ | 0,327 | -1,27 |
| $V^{2+} + 2e^- = V$ | -1,175 | — |
| $V^{3+} + e^- = V^{2+}$ | -0,255 | — |
| $VO^{2+} + 2H^+ + e^- = V^{2+} + H_2O$ | 0,337 | — |
| $WO_2 + 4H^+ + 4e^- = W + 2H_2O$ | -0,119 | — |
| $WO_3 + 6H^+ + 6e^- = W + 3H_2O$ | -0,090 | -0,40 |
| $Y^{3+} + 3e^- = Y$ | -2,372 | 0,18 |
| $Y(OH)_3 + 3e^- = Y + 3OH^-$ | -2,81 | -0,95 |
| $Zn^{2+} + 2e^- = Zn$ | -0,761 8 | 0,091 |
| $Zn(OH)_2 + 2e^- = Zn + 2OH^-$ | -1,245 | -1,002 |
| $ZnO_2^{2-} + 2H_2O + 2e^- = Zn + 4OH^-$ | -1,215 | — |
| $ZnSO_4 + 2e^- = Zn(Hg) + SO_4^{2-}$ | -0,799 3 | — |
| $Zr^{4+} + 4e^- = Zr$ | -1,529 | — |
| $H_2ZrO_3 + H_2O + 4e^- = Zr + 4OH^-$ | -2,36 | -1,11 |
| $ZrO_2 + 4H^+ + 4e^- = Zr + 2H_2O$ | -1,553 | — |

8.16 ELEKTRODOVÉ POTENCIÁLY NĚKTERÝCH REFERENTNÍCH ELEKTROD PŘI RŮZNÝCH TEPLOTÁCH

Elektroda kalomelová

| Elektrolyt | $t/^\circ C$ | | | | |
|------------|--------------|---------|---------|---------|---------|
| | 10 | 15 | 20 | 25 | 30 |
| 0,1M-KCl | 0,336 2 | 0,336 1 | 0,335 8 | 0,335 6 | 0,335 4 |
| 1,0M-KCl | 0,286 8 | 0,285 2 | 0,284 4 | 0,283 0 | 0,281 5 |
| nas. KCl | 0,253 9 | 0,251 1 | 0,247 8 | 0,244 5 | 0,241 2 |

Elektroda argentchloridová

| Elektrolyt | $t/^\circ C$ | | | | |
|------------|--------------|---------|---------|---------|-------|
| | 10 | 15 | 20 | 25 | 30 |
| 0,1M-KCl | — | — | — | 0,288 | — |
| 1,0M-KCl | 0,231 4 | 0,228 6 | 0,225 6 | 0,222 3 | 0,219 |
| nas. KCl | — | — | — | 0,198 | — |

Elektroda merkurosulfátová

| Elektrolyt | $t/^\circ C$ | | | | |
|-------------------------------------|--------------|---------|---------|---------|---------|
| | 10 | 15 | 20 | 25 | 30 |
| H_2SO_4 ($a_{SO_4^{2-}} = 1$) | 0,627 0 | 0,623 1 | 0,619 3 | 0,615 2 | 0,611 1 |
| 0,5M-K ₂ SO ₄ | — | — | — | 0,682 | — |
| nas. K ₂ SO ₄ | — | — | — | 0,650 | — |

Elektroda merkurioxidová

| Elektrolyt | $t/^\circ C$ | | | | |
|------------|--------------|----|----|-------|----|
| | 10 | 15 | 20 | 25 | 30 |
| 0,1M-NaOH | — | — | — | 0,165 | — |
| 1,0M-NaOH | — | — | — | 0,140 | — |