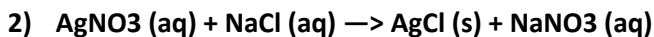


NAME: Lía Prieto García

SAMPLE: Sodium chloride

1) IDENTIFICATION REACTIONS OF IONS

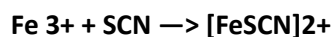


- CATIONS (*describe briefly reactions*):

$\text{Na} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 (\text{aq})$ no precipitate

$\text{Na} + \text{HNO}_3 \rightarrow \text{NaNO}_3 (\text{aq})$ no precipitate

$\text{Ag}^+ (\text{waste}) + \text{SCN}^- \rightarrow \text{AgSCN} (\text{s})$ The excess of thiocyanate reacts with Fe^{3+} to form a dark red complex, When the silver ions have reacted.



- ANIONS (*describe briefly reactions*):

$\text{Cl}^- + \text{AgNO}_3 \rightarrow \text{AgCl} (\text{s})$ precipitate

2) ASSAY

Volumetric solutions: 0'1M silver nitrate

Titre of volumetric solutions: 0'9998

Titration No.	m [g] (<i>4 decimal places</i>)	Consumption of VS [ml]	ASSAY
1.	0.9968	8.21	98.8345
2.	0.9954	8.12	99.4965
3.	1.0045	8.45	96.6960
4.	1.0354	7.89	97.0027
Average			98.0074

CALCULATION PROCEDURE:

STATISTICAL EVALUATION:**Range:****R = 2.8005****Standard deviation** (*estimated from range*):**sd = 1.3602****Relative standard deviation:****RSD = 1.3879**

CONCLUSION (*does your sample meet/not meet Ph. Eur*): **The sample do not meet Ph. Eur, because the value 98.0074 is not in the interval 99.0-100.5.**