

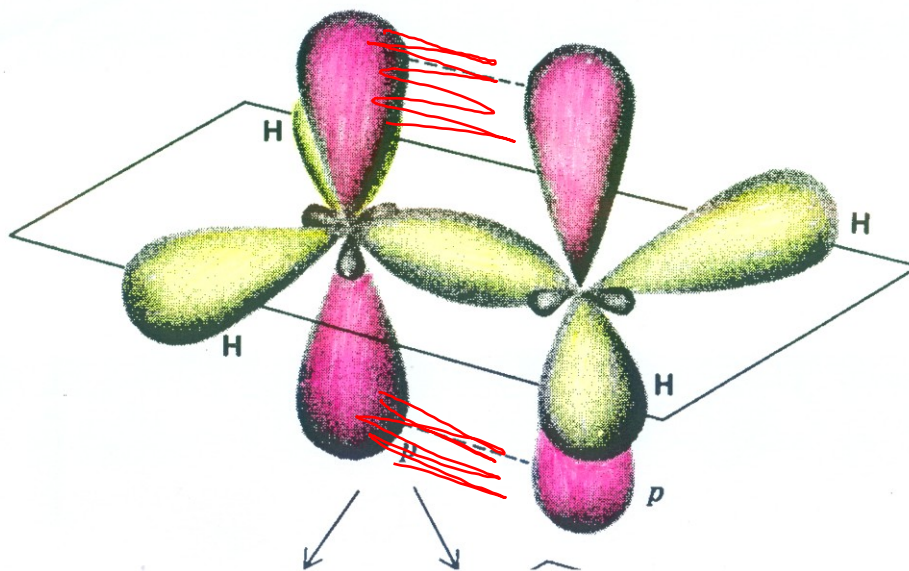
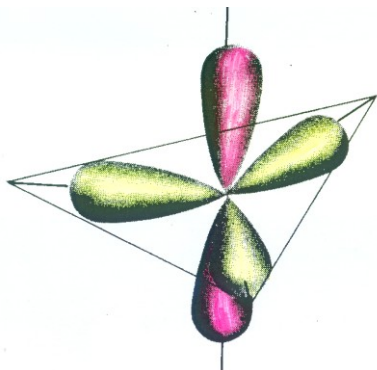
# Alkeny a cykloalkeny

## Alkeny a cykloalkeny

Sloučeniny s vazbami **C - C** 0,154 nm a vazby **C - H**  
**C = C** 0,134 nm a vazby **=C-H**

**Energie vazeb**

**C - C** 351 kJ/mol  $\sigma$  - vazba  
**C = C** 610 kJ/mol  $\Delta = 259$  kJ/mol  $\pi$  - vazba



# Alkeny a cykloalkeny

## Názvosloví alkenů

$\text{H}_2\text{C}=\text{CH}_2$  ethen, ethylen

$\text{H}_2\text{C}=\text{CH}-\text{CH}_3$  propen, propylen

$\text{H}_2\text{C}=\text{CH}-\text{CH}_2-\text{CH}_3$  buten

$\text{H}_3\text{C}-\underset{\text{H}}{\text{C}}=\text{CH}-\text{CH}_3$  but-2-en

$\text{H}_2\text{C}=\text{CH}-$

$\text{H}_2\text{C}=\text{CH}-\text{CH}_2-$

$\text{H}_3\text{C}-\text{CH}=\text{CH}-$

$\text{H}_2\text{C}=\underset{|}{\text{C}}-\text{CH}_3$

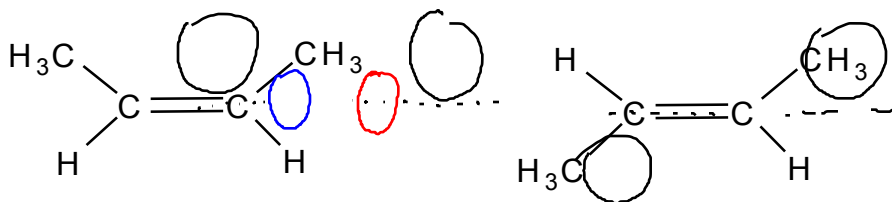
ethenyl- , **vinyl-**

prop-2-en-1-yl, **allyl-**

prop-1-en-1-yl

prop-1-en-2-yl

## Geometrická isomerie



cis-but-2-en

(Z)-but-2-en

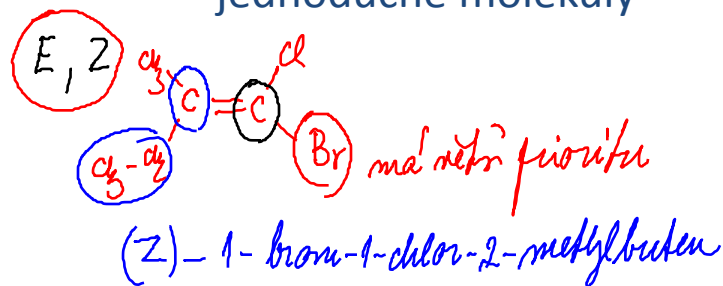
trans-but-2-en

(E)-but-2-en

(E)

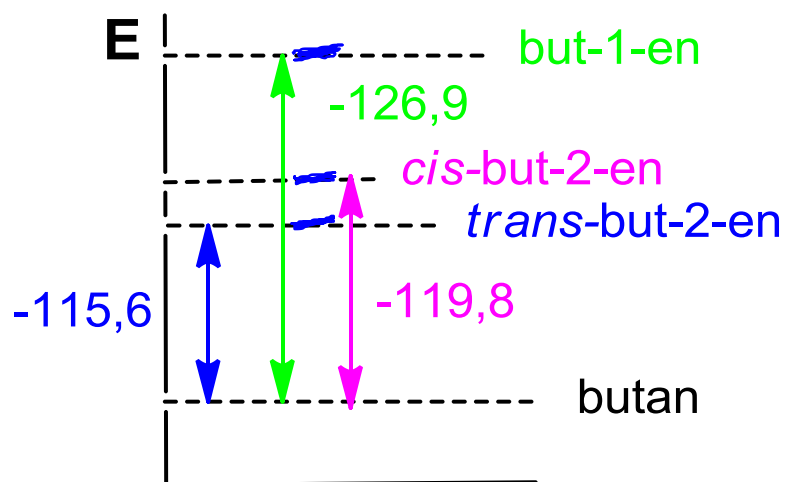
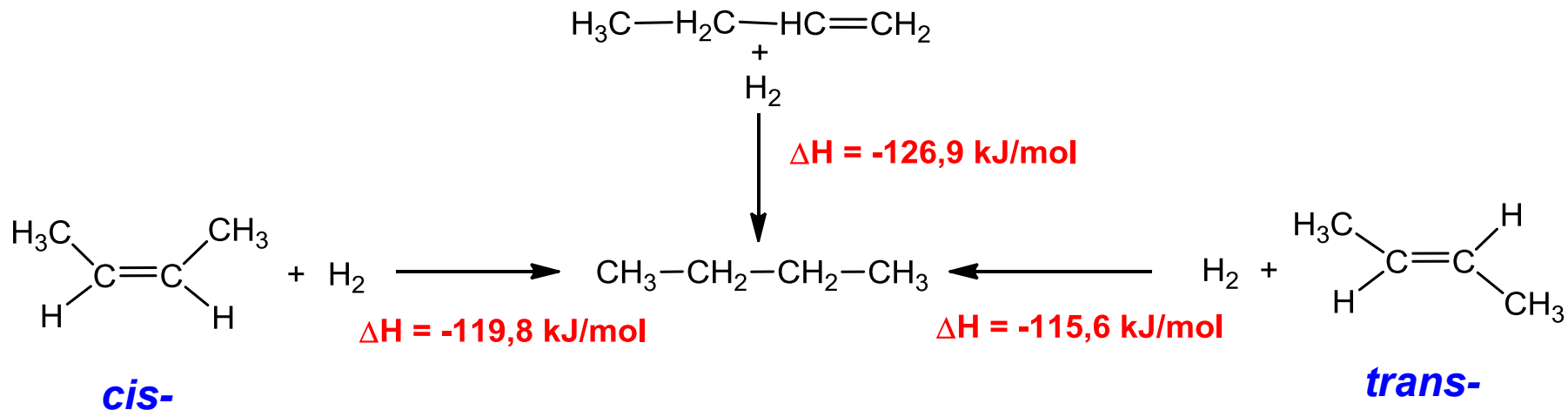
(Z)

používání afixů **cis-** a **trans-** vhodné pro jednoduché molekuly



# Alkeny a cykloalkeny

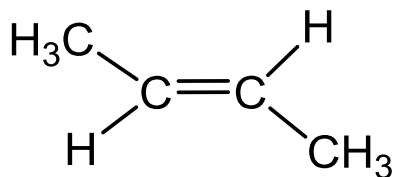
## STABILITA alkenů



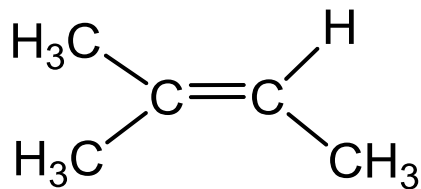
*hydrogenační teplo*

**stabilita *trans*-isomeru je větší**

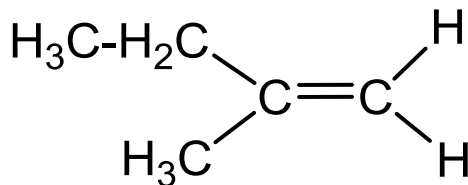
## Alkeny a cykloalkeny



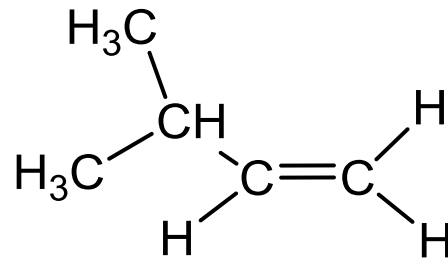
$$\Delta H = -115,6 \text{ kJ/mol}$$



$$\Delta H = -112,7 \text{ kJ/mol}$$



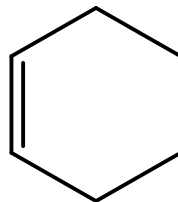
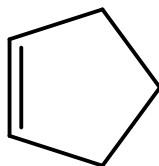
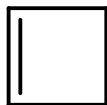
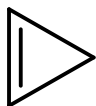
$$\Delta H = -119,4 \text{ kJ/mol}$$



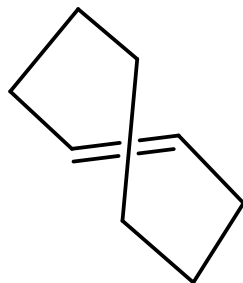
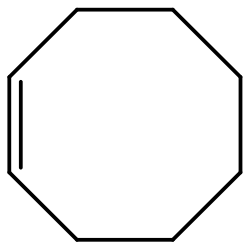
$$\Delta H = -126,9 \text{ kJ/mol}$$

čím víc je systém rozvětvený na násobné vazbě, tím je stabilnější

## Alkeny a cykloalkeny



všechny cykloalkeny do 6 atomů v kruhu jsou známy jako *cis*- isomery



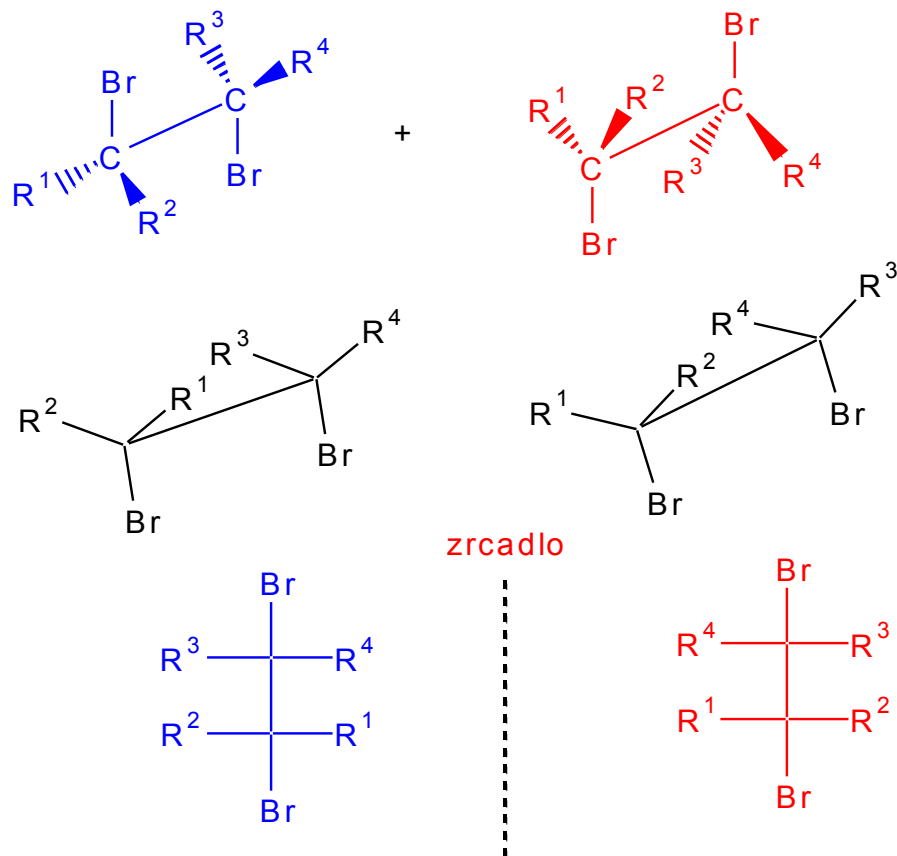
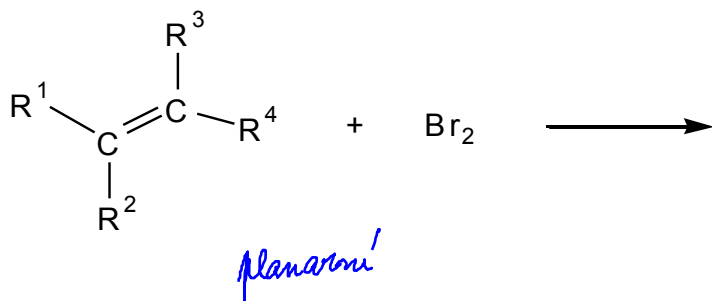
stálý a izolovatelný je až *trans*-cykloookten

# Alkeny a cykloalkeny

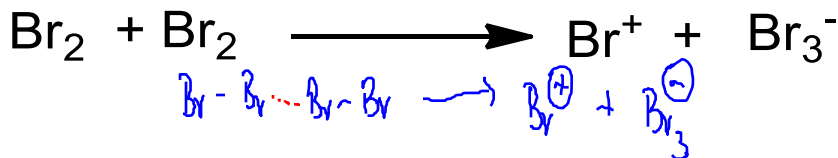
## HLAVNÍ SMĚR REAKTIVITY ALKENŮ

## ADIČNÍ REAKCE

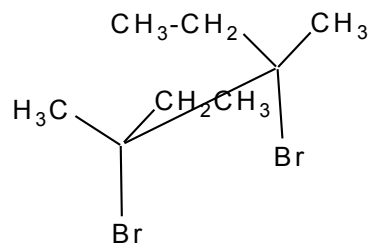
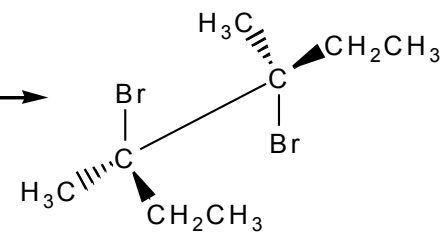
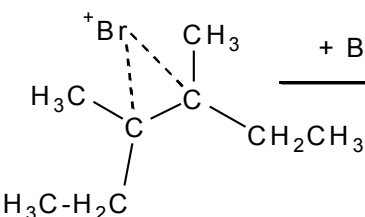
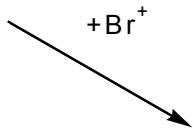
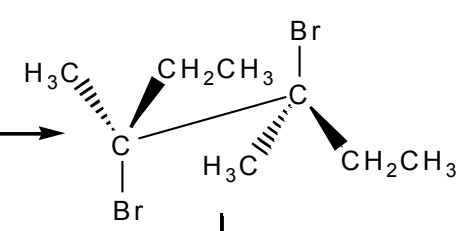
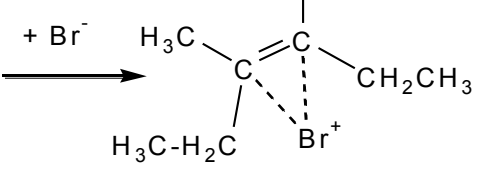
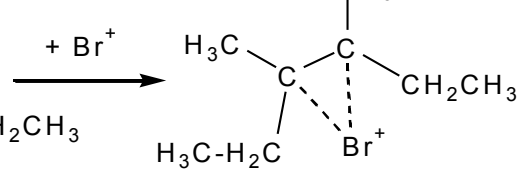
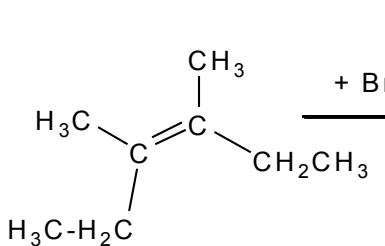
Reakce probíhají stereospecificky



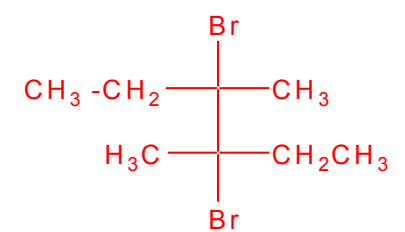
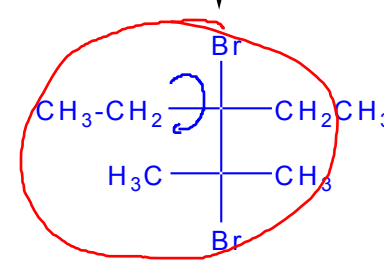
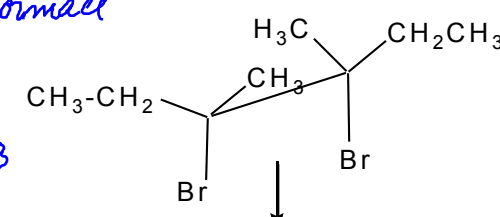
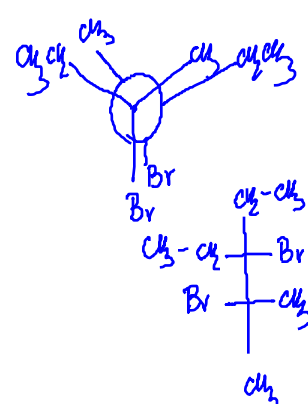
Při adici symetrické molekuly na nesymetricky substituovaný olefin vznikají enantiomery – **reakce je stereospecifická trans- (anti-)**



# Stereospecifický průběh adičních reakcí

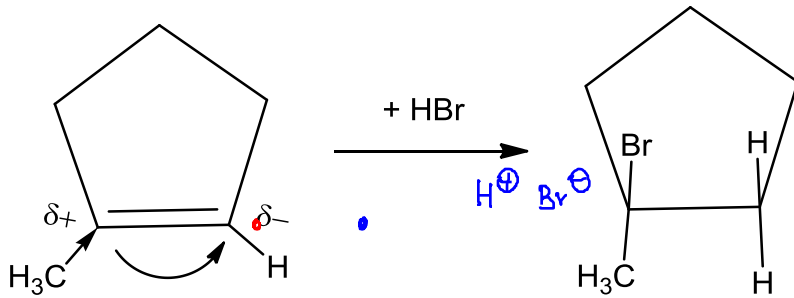


*rotovaná konformace*



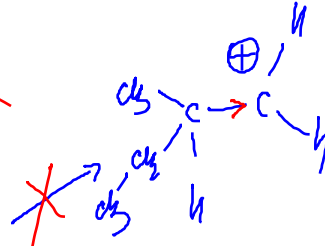
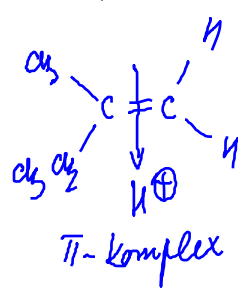
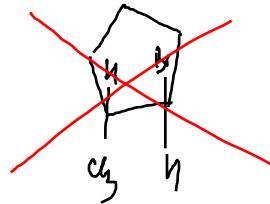
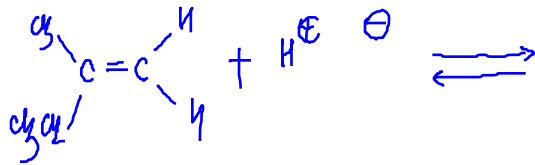


# Regiospecifický průběh adičních reakcí

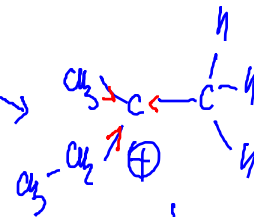


*nesymetrická molekula, které se aduje*  
*nesym. uel molekula, které se aduje*

Markovnikovovo pravidlo

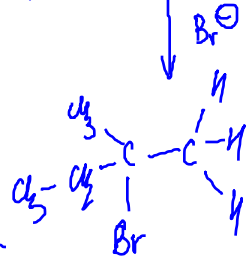
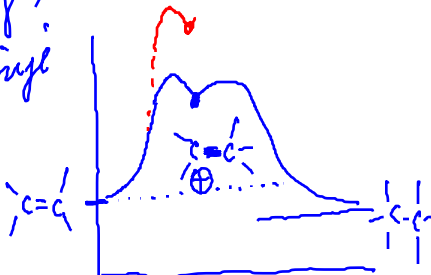


*primární karbokatión*

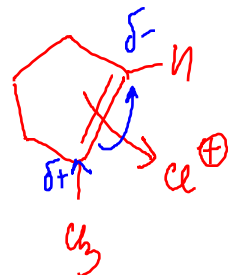
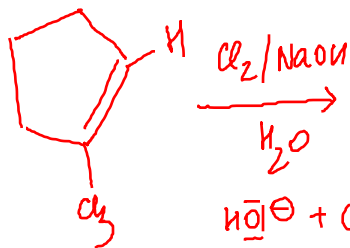
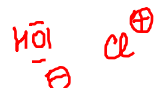
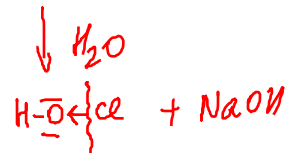
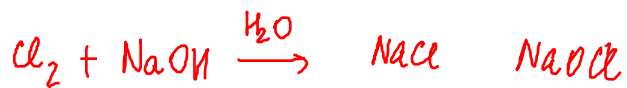


*terc. karbokatión*

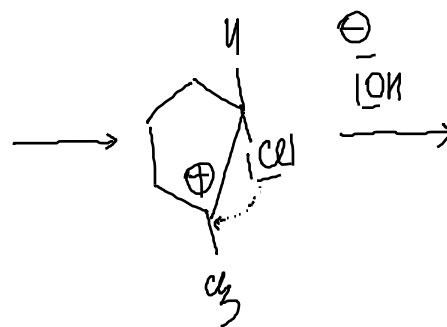
- elektrofilní částice se připojuje na silnější a více vodlivý
- elektrofilní částice se připojuje k uhlíku, kde aby vznikl stabilnější karbokatión



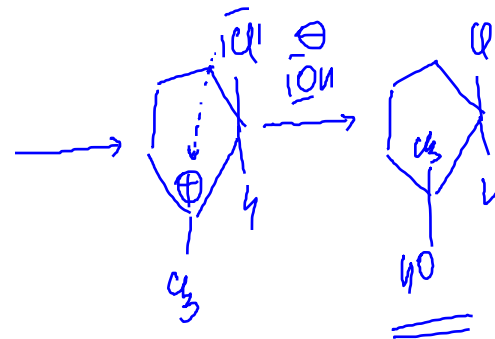
# Regiospecifický průběh adičních reakcí



$\pi$ -komplex



(1S,2S) -  
-chlor-1-methylcyclohexan



# Regiospecifický průběh adičních reakcí

