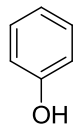
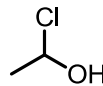
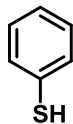


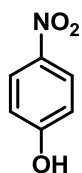
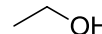
1. V uvedených dvojicích označte, která molekula bude silnější kyselina.



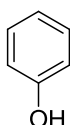
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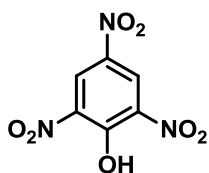
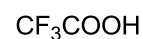
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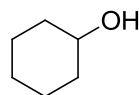
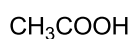
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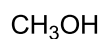
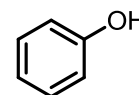
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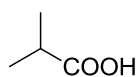
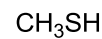
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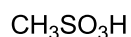
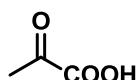
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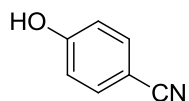
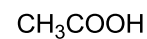
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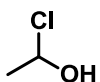
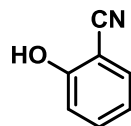
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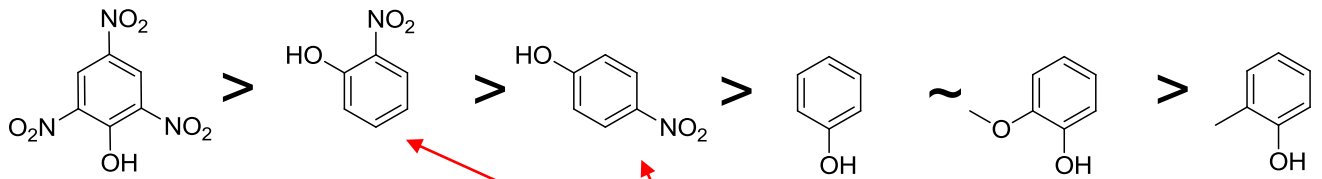


PAMATOVAT!

- sirmé deriváty jsou kyselejší než kyslíkaté (thioly > alkoholy; thiofenoly > fenoly)
- akceptorní skupiny zvyšují kyselost a snižují bazicitu
- donorní skupiny snižují kyselost a zvyšují bazicitu

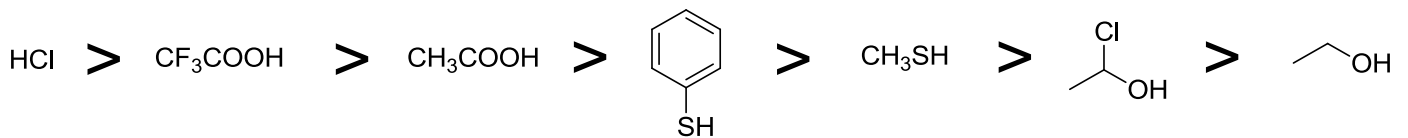
2. Seřadte sloučeniny podle klesající kyselosti

A.

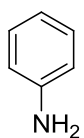


akceptorní skupina v ortho-poloze bude mít na kyselost OH skupiny větší vliv z důvodu menší vzdálenosti než akceptor v para-poloze

B.

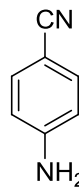


3. V uvedených dvojicích označte, která molekula bude silnější báze.

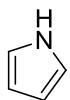
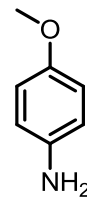


<

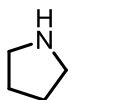
NH₃



<

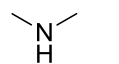


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NH₃

<



PAMATOVAT!

- alifatické aminy jsou bazičtější než aromatické
- alifatické aminy jsou bazičtější než amoniak (NH₃) díky donornímu efektu alkyl. skupin
- ve skutečnosti bazicita roste ve směru: aromatické aminy < NH₃ < 1° ~ 3° < 2°
- donorní skupiny zvyšují bazicitu a snižují kyselost
- akceptorní skupiny snižují bazicitu a zvyšují kyselost

4. Seřad'te podle klesající bazicity

