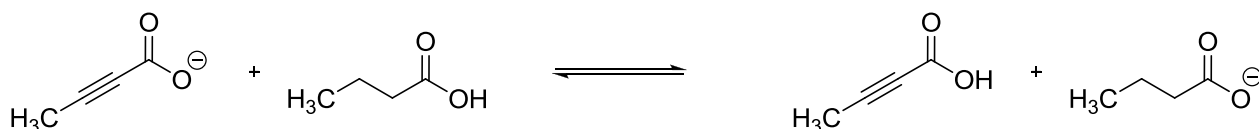
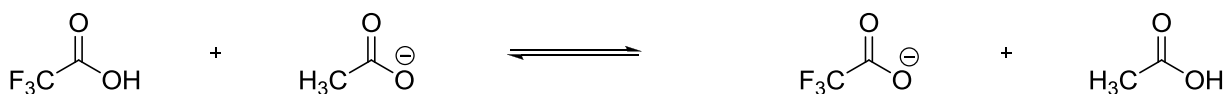
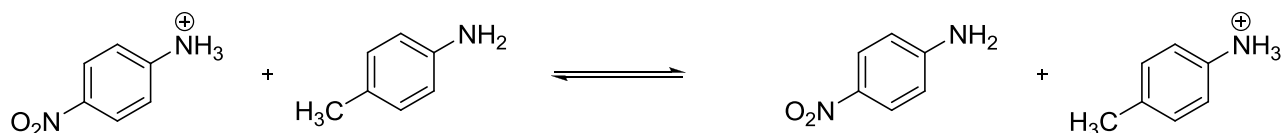
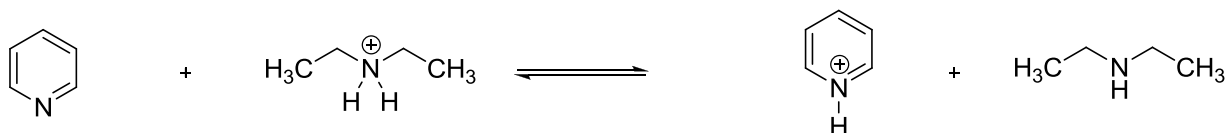
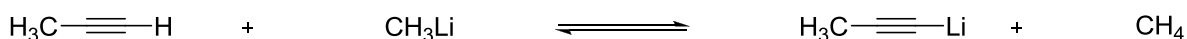
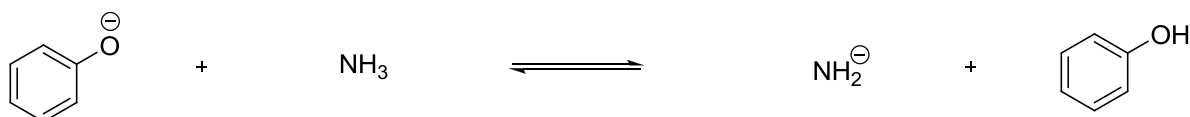
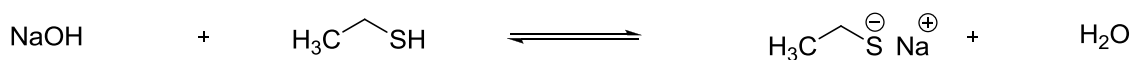
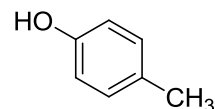
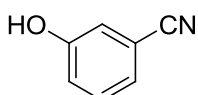
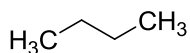
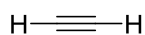
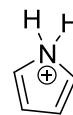
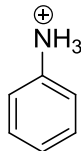
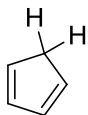
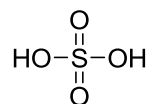
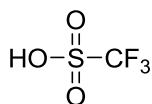
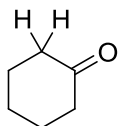
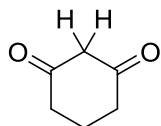


- Nižší  $pK_a \rightarrow$  silnější kyselina
- Acidobazická rovnováha je posunuta na stranu slabší kyseliny
- Akceptorní skupiny zvyšují kyselost a snižují bazicitu
- Donorní skupiny zvyšují bazicitu a snižují kyselost
- Čím silnější kyselina, tím slabší je její konjugovaná báze
- Čím slabší kyselina, tím silnější je její konjugovaná báze
- Sílu kyseliny můžeme odhadovat vyhodnocením stabilizačních efektů její konjugované báze  $\rightarrow$  čím více bude báze stabilizovaná, tím ochotněji bude kyselina disociovat (tím silnější bude)
- Tvrdé částice mají malý poloměr a malou polarizovatelnost
- Měkké částice mají větší poloměr a velkou polarizovatelnost

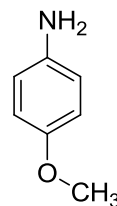
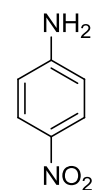
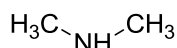
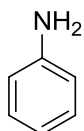
Určete, kam bude posunuta acidobazická rovnováha.



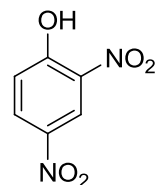
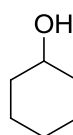
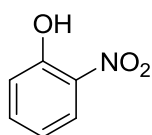
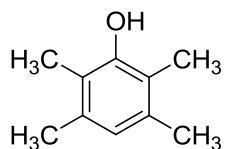
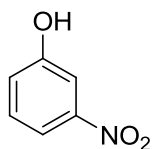
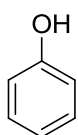
V uvedených dvojicích označte silnější kyselinu:



Seřad'te dle stoupající bazicity:



Seřad'te dle stoupající kyselosti:



Ve dvojici označte měkčí kyselinu:



Ve dvojici označte měkčí bázi:

