

## LESSON 8: ACIDS, BASES AND SALTS

### I. What is the pH scale? What does it measure?

#### II.

- Put the strips of papers with food items along the scale, according to their pH factor, from high alkaline, to alkaline, low alkaline, low acid, acid to high acid.
- What problems might consuming too many acidic foods cause?

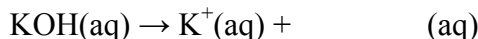
### III. Acids, bases and salts – definitions; watch the video and complete the text below with missing pieces of information. (source: [www.gpb.org/chemistry-physics/chemistry/1101](http://www.gpb.org/chemistry-physics/chemistry/1101), visited on October 18, 2011)

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behaviour of acids and bases.

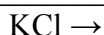
An Arrhenius acid is defined as any compound that \_\_\_\_\_ in aqueous solution to form \_\_\_\_\_ ions.



An Arrhenius base is defined as any compound that \_\_\_\_\_ in aqueous solution to form \_\_\_\_\_ ions.



Salts are compounds that \_\_\_\_\_ in aqueous solution releasing \_\_\_\_\_ ions.



### IV. Classifying compounds; using the Arrhenius definition, classify the following examples as acids, bases and salts (source: [www.gpb.org/chemistry-physics/chemistry/1101](http://www.gpb.org/chemistry-physics/chemistry/1101), visited on October 18, 2011)

HBr

Mg(OH)<sub>2</sub>

HCl

KNO<sub>2</sub>

HFO<sub>4</sub>

Ba(OH)<sub>2</sub>

KCl

H<sub>3</sub>PO<sub>4</sub>

HClO

Al(OH)<sub>3</sub>

KC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>

NaCl

### V. Naming acids and bases; (source: [www.gpb.org/chemistry-physics/chemistry/1101](http://www.gpb.org/chemistry-physics/chemistry/1101), visited on October 18, 2011)

a) watch the video and complete the text below

Since bases are \_\_\_\_\_ compounds, they are named in the usual way:

NH<sub>4</sub>OH –

Al(OH)<sub>3</sub> –

Binary acids consist of \_\_\_\_\_ elements, the first being \_\_\_\_\_. Binary acids are named using the format: \_\_\_\_\_+(root word of second element)+IC acid

Ternary acids consist of \_\_\_\_\_ elements. Do not use a prefix. Simply change the ending of the polyatomic ion's name and add the word \_\_\_\_\_. -ate ending becomes \_\_\_\_\_ and -ite becomes \_\_\_\_\_.

b) Now name the following acids:

HBr	H <sub>3</sub> PO <sub>3</sub>
HNO <sub>3</sub>	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>
HNO <sub>2</sub>	H <sub>2</sub> CO <sub>3</sub>
HI	HClO <sub>2</sub>
	HF
	H <sub>2</sub> SO <sub>3</sub>

## VI. Naming salts;

a) read the text below, and according to the information given, name the salts underneath, as in the example. ([http://www.bbc.co.uk/schools/gcsebitesize/science/add\\_aqa/ions/acidsbasesrev4.shtml](http://www.bbc.co.uk/schools/gcsebitesize/science/add_aqa/ions/acidsbasesrev4.shtml) , accessed 18th October, 2011)

The name of a salt has two parts. The first part comes from the metal in the base or carbonate, or the metal itself if a reactive metal like magnesium or zinc is used. The second part of the name comes from the acid used to make it. The names of salts made from hydrochloric acids end in **-chloride**, while the name of salts made from sulfuric acid end in **-sulfate**.

<b>metal</b>		<b>acid</b>		<b>salt</b>
<b>sodium</b> hydroxide	reacts with	hydro <b>chloric</b> acid	to make	<b>sodium chloride</b>
1. <b>copper</b> oxide		hydro <b>chloric</b> acid		
2. <b>sodium</b> hydroxide		<b>sulfuric</b> acid		
3. <b>zinc</b> oxide		<b>sulfuric</b> acid		
4. <b>ammonia</b>		hydro <b>chloric</b> acid		

b) write formulas of the reactions above

- 1.
- 2.
- 3.
- 4.

## VII. Give formulas of these acids, bases and salts

boron silicide	magnesium phosphide
sodium hydroxide	zinc hydroxide
iron(III)chloride	aluminium sulfide
sulfuric acid	sulfurous acid

## VIII. Chemistry quiz

### Test Bite

[http://www.bbc.co.uk/apps/ipl/schools/gcsebitesize/science/quizengine?quiz=add\\_aqa\\_acidsbasestest&templateStyle=science](http://www.bbc.co.uk/apps/ipl/schools/gcsebitesize/science/quizengine?quiz=add_aqa_acidsbasestest&templateStyle=science)

1. A solution has a pH of 4 - what does this mean?
  - It is acidic.
  - It is neutral.
  - It is alkaline.
2. Which of the statements below is correct?
  - Bases are acids that dissolve in water.
  - Bases are alkalis that dissolve in water.
  - Alkalis are bases that dissolve in water.
3. A liquid has a pH of 7. What does this tell you about the liquid?
  - It is water.
  - It is sodium chloride solution.
  - It is neutral.
4. Which salt is made when calcium carbonate reacts with hydrochloric acid?
  - sodium chloride
  - calcium chloride
  - calcium sulphate
5. Which pair of substances will react together to make copper sulfate?
  - copper and sulfuric acid
  - copper oxide and sulfuric acid
  - copper oxide and hydrochloric acid
6. Which is the correct order of methods for making a salt from an acid and an insoluble base?
  - filtration ==> evaporation ==> neutralisation
  - neutralisation ==> evaporation ==> filtration
  - neutralisation ==> filtration ==> evaporation

## GRAMMAR REVISION: Defining relative clauses

- I. complete the sentences below with an appropriate relative pronoun: (1&2: Oreska et al, 2006, *English for Chemists*)
1. Elements are substances \_\_\_\_\_ cannot be broken down by chemical methods any further.
  2. Silver and gold are elements \_\_\_\_\_ are widely used in jewellery.
  3. Kampus is the part of Masaryk University \_\_\_\_\_ seminar rooms are equipped with modern audio-visual systems.
  4. Salt \_\_\_\_\_ comes from the sea is considered to be the best for cooking.
  5. Our teacher is the person \_\_\_\_\_ instructions we must obey.
  6. The beaker \_\_\_\_\_ you use for experiments has to be cleaned afterwards.
  7. The scientists \_\_\_\_\_ discovered the presence of sodium in the Sun are Robert Bunsen and Gustav Kirchhoff.
  8. \_\_\_\_\_ of you can describe a Bunsen burner?
  9. \_\_\_\_\_ knows the symbols of all the chemical elements?
  10. Destructive distillation is a method \_\_\_\_\_ involves separating a mixture of several components of different boiling points.
- II. Make definitions of the terms given to you by the teacher

## VOCABULARY

ENGLISH	CZECH
acid	kyselina, kyselý, acidní
alkaline	zásaditý, alkalický
aqueous solution	vodný roztok
behaviour	chování, reakce, reagování
binary	dvojkový, binární
carbonate	uhličitan
dissociate	odloučit se, disociovat
ion	iont
ionization	ionizace
polyatomic	viceatomový
release	uvolnit
ternary	trojitý, ternární

