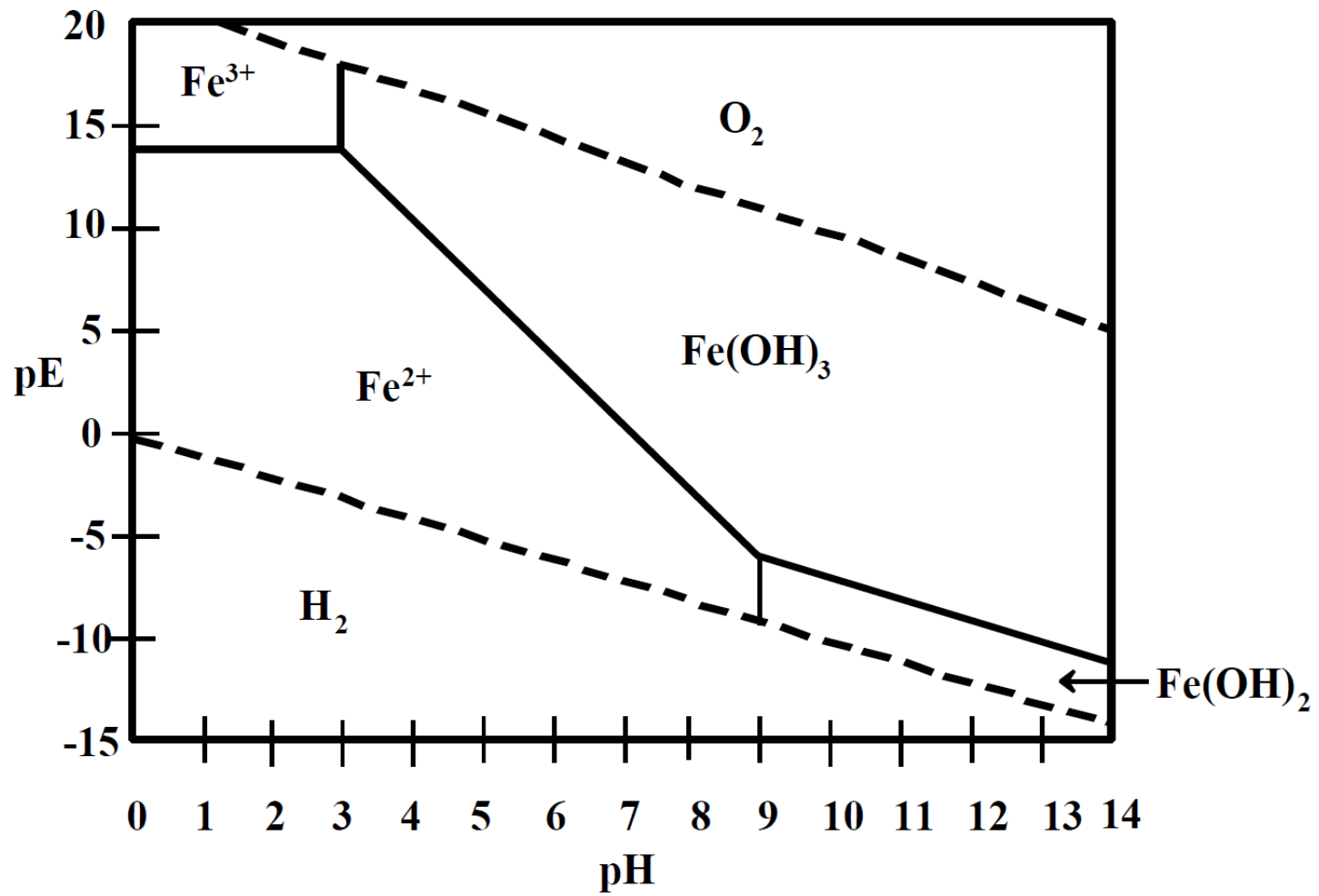


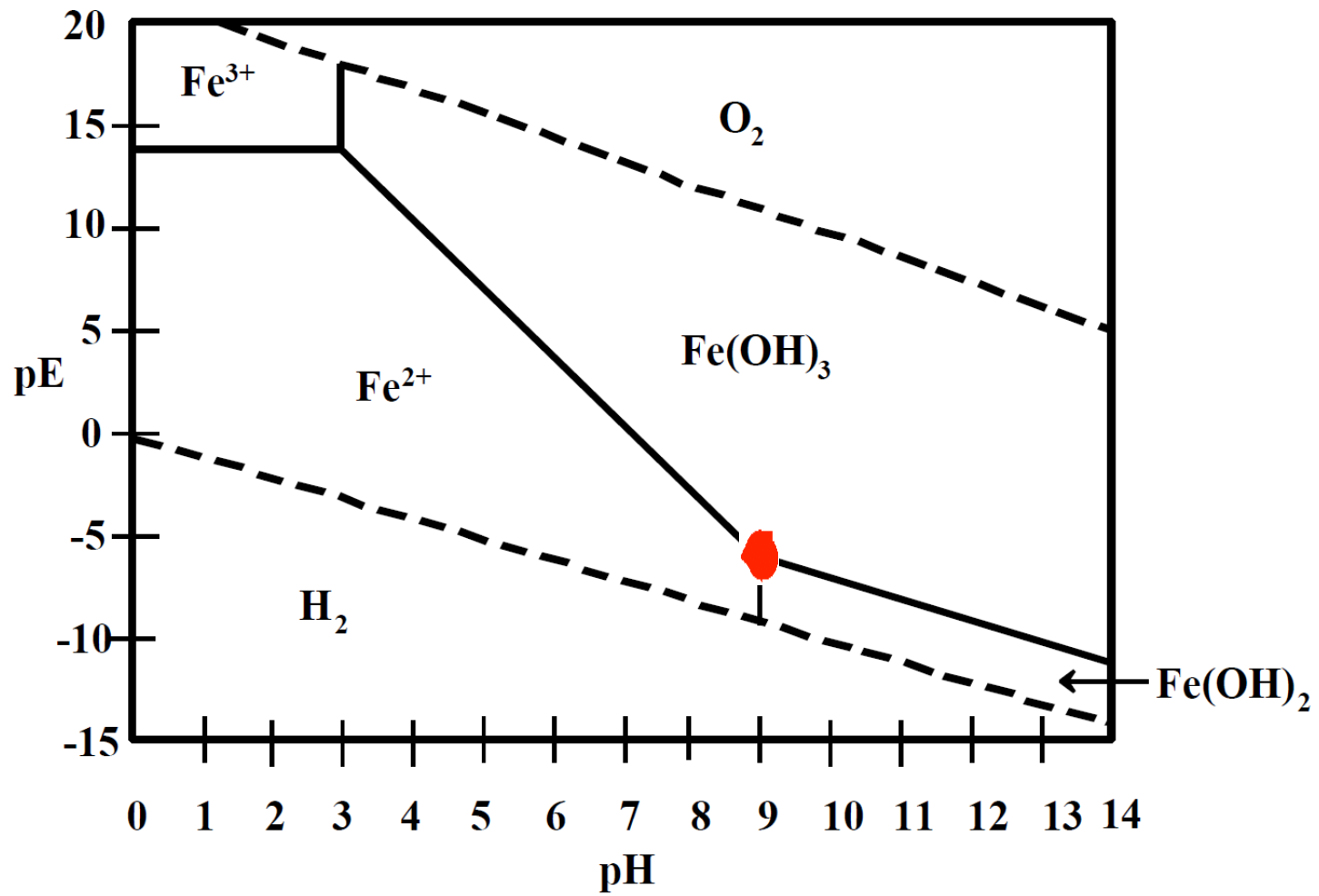
Vypočítejte hodnoty $[\text{Fe}^{3+}]$, pE a pH v bodě na grafu pE-pH, kde Fe^{2+} při koncentraci $1,00 \times 10^{-5} \text{ M}$, $\text{Fe}(\text{OH})_2$ a $\text{Fe}(\text{OH})_3$ jsou všechny v rovnováze.

Součin rozpustnosti

$$\text{Fe}(\text{OH})_2 = 8,0 \times 10^{12}$$

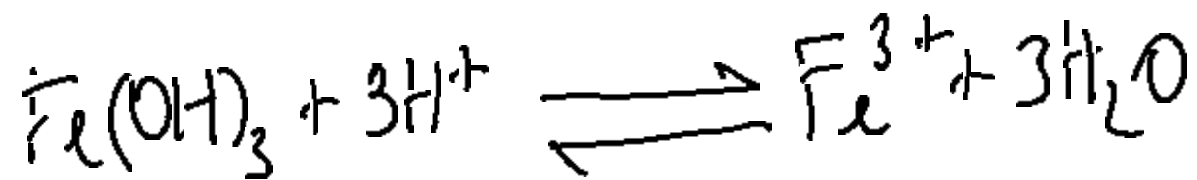
$$\text{Fe}(\text{OH})_3 = 9,1 \times 10^3$$





$$pH = 8.95$$

$$pE = -4.65$$



$$K_{sp} = \frac{[\text{Fe}^{3+}]}{[\text{H}^+]^3} = 9.1 \cdot 10^3$$

$$[\text{Fe}^{3+}] = 9.1 \cdot 10^3 \cdot [10^{-8.95}]^3 =$$